

Write your name here

Surname

Other names

Centre Number

Candidate Number

**Edexcel GCE****Chemistry****Advanced****Unit 5: General Principles of Chemistry II – Transition Metals and Organic Nitrogen Chemistry (including synoptic assessment)**

Tuesday 19 June 2012 – Afternoon

**Time: 1 hour 40 minutes**

Paper Reference

**6CH05/01****You must have: Data Booklet**

Total Marks

**Candidates may use a calculator.****Instructions**

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need.*

**Information**

- The total mark for this paper is 90.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (\*) are ones where the quality of your written communication will be assessed – *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*
- A Periodic Table is printed on the back cover of this paper.

**Advice**

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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**PEARSON**

## SECTION A

Answer ALL the questions in this section. You should aim to spend no more than 20 minutes on this section. For each question, select one answer from A to D and put a cross in the box . If you change your mind, put a line through the box  and then mark your new answer with a cross .

1 The acid used in a standard hydrogen electrode to provide a  $1 \text{ mol dm}^{-3}$  solution of hydrogen ions is

- A ethanoic acid.
- B phosphoric(V) acid.
- C sulfuric acid.
- D hydrochloric acid.

(Total for Question 1 = 1 mark)

2 To measure the standard electrode potential for the  $\text{Ag}^+(\text{aq})|\text{Ag}(\text{s})$  electrode, the most suitable chemical for the solution in a salt bridge to connect the two half cells is

- A potassium chloride.
- B potassium iodide.
- C potassium nitrate.
- D potassium sulfate.

(Total for Question 2 = 1 mark)

3 For any reversible reaction,  $\ln K$  is proportional to

- A both  $\Delta S_{\text{total}}$  and  $E_{\text{cell}}$ .
- B  $\Delta S_{\text{total}}$  but not  $E_{\text{cell}}$ .
- C  $E_{\text{cell}}$  but not  $\Delta S_{\text{total}}$ .
- D neither  $E_{\text{cell}}$  nor  $\Delta S_{\text{total}}$ .

(Total for Question 3 = 1 mark)



- 4 This question is about a titration to determine the iron content of a tablet. The iron(II) ions in the tablet are oxidized to iron(III) ions by acidified manganate(VII) ions which are reduced to manganese(II) ions.

(a) The mole ratio of iron(II) to manganate(VII) ions in the reaction is

(1)

	$\text{Fe}^{2+}$	$\text{MnO}_4^-$
<input type="checkbox"/> A	1	5
<input type="checkbox"/> B	2	5
<input type="checkbox"/> C	5	2
<input type="checkbox"/> D	5	1

- (b) A 0.200 g tablet is dissolved to make exactly 100 cm<sup>3</sup> of solution. 10 cm<sup>3</sup> of this solution is found to contain  $5.38 \times 10^{-5}$  mol of iron(II) ions.

The percentage by mass of iron ( $A_r = 55.8$ ) in the tablet is

(1)

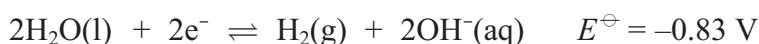
- A  $\frac{5.38 \times 10^{-5} \times 55.8}{0.200} \times \frac{100}{10} \times 100\%$
- B  $\frac{5.38 \times 10^{-5}}{55.8 \times 0.200} \times \frac{100}{10} \times 100\%$
- C  $\frac{5.38 \times 10^{-5} \times 55.8}{0.200} \times \frac{10}{100} \times 100\%$
- D  $\frac{5.38 \times 10^{-5} \times 0.200}{55.8} \times \frac{10}{100} \times 100\%$

(Total for Question 4 = 2 marks)

Use this space for any rough working. Anything you write in this space will gain no credit.



- 5 Some fuel cells are used to produce energy from the combination of hydrogen and oxygen. The half-equations for a hydrogen-oxygen fuel cell in alkaline solution are



(a) The equation for the cell reaction is

(1)

- A  $3\text{H}_2\text{O}(\text{l}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2(\text{g}) + 4\text{OH}^-(\text{aq})$
- B  $\text{H}_2(\text{g}) + 4\text{OH}^-(\text{aq}) \rightarrow 3\text{H}_2\text{O}(\text{l}) + \frac{1}{2}\text{O}_2(\text{g})$
- C  $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$
- D  $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$

(b) From the data above,  $E_{\text{cell}}$  for the reaction in the hydrogen-oxygen fuel cell is

(1)

- A  $-0.43 \text{ V}$
- B  $+0.43 \text{ V}$
- C  $-1.23 \text{ V}$
- D  $+1.23 \text{ V}$

(c) An advantage of operating a hydrogen fuel cell over a conventional jet engine for powering an aeroplane is

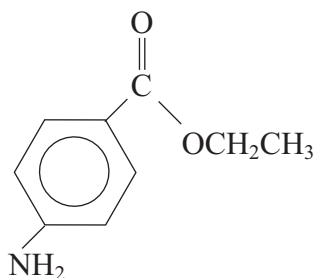
(1)

- A it is 100% efficient.
- B it does not produce greenhouse gas emissions.
- C it does not produce emissions which damage the ozone layer.
- D hydrogen is easier to transport.

(Total for Question 5 = 3 marks)



6 Benzocaine is used as a local anaesthetic.



Separate samples of a solution of benzocaine are added to 2,4-dinitrophenylhydrazine, hot aqueous sodium hydroxide, and dilute hydrochloric acid.

Which chemicals react with benzocaine?

- A All three
- B Only sodium hydroxide and hydrochloric acid
- C Only hydrochloric acid
- D Only sodium hydroxide

(Total for Question 6 = 1 mark)

7 In the first commonly used breathalyser, acidified potassium dichromate(VI) was used which was reduced to chromium(III) by alcohol.

The colour change seen when alcohol was present in a motorist's breath is from

- A orange to green.
- B orange to yellow.
- C yellow to green.
- D yellow to orange.

(Total for Question 7 = 1 mark)

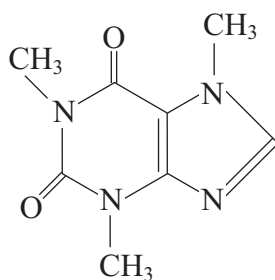
8 For which of the following are both tests acceptable evidence for a driver to be successfully prosecuted for excess alcohol in the blood in many countries?

- A A dichromate(VI) breath analyser and a blood test.
- B A dichromate(VI) breath analyser and a fuel cell breath analyser.
- C A dichromate(VI) breath analyser and an infrared breath analyser.
- D An infrared breath analyser and a blood test.

(Total for Question 8 = 1 mark)



9 The structural formula of caffeine,  $C_8H_{10}O_2N_4$ , is shown below.



(a) How many main peaks would you expect in the proton nuclear magnetic resonance spectrum of caffeine?

(1)

- A 1
- B 2
- C 3
- D 4

(b) At which of the following wavenumbers is an absorption peak **not** present in the infrared spectrum of caffeine?

(1)

- A  $3600\text{ cm}^{-1}$
- B  $2925\text{ cm}^{-1}$
- C  $1690\text{ cm}^{-1}$
- D  $1660\text{ cm}^{-1}$

(c) The parent ion peak of caffeine in the mass spectrum of caffeine would be at  $m/e$  ratio

(1)

- A 101
- B 102
- C 193
- D 194



(d) Intermolecular forces between caffeine molecules would include

(1)

- A London forces only.
- B London forces and hydrogen bonds.
- C London forces and permanent dipole forces.
- D London forces, permanent dipole forces, and hydrogen bonds.

(Total for Question 9 = 4 marks)

10 (a) The principal advantage of combinatorial chemistry over traditional methods for developing pharmaceuticals is that

(1)

- A very small amounts of compounds are used.
- B many more compounds can be made in a given time.
- C compounds formed are more powerful drugs.
- D compounds formed have fewer side effects.

(b) A refinement of combinatorial chemistry involves initially attaching compounds firmly to polymer beads. The type of attachment is most likely to be by

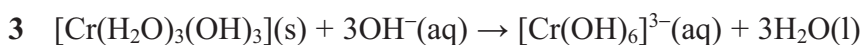
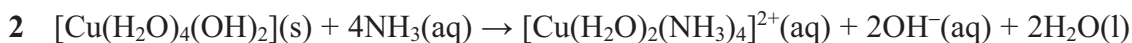
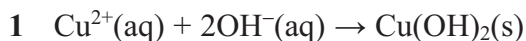
(1)

- A metallic bonds.
- B covalent bonds.
- C London forces.
- D permanent dipole forces.

(Total for Question 10 = 2 marks)



11 Four reactions involving the transition elements copper and chromium are given below.



(a) Which reaction produces a dark blue solution?

(1)

A 1

B 2

C 3

D 4

(b) Which two reactions show the amphoteric behaviour of a metal hydroxide?

(1)

A 1 and 2

B 2 and 3

C 2 and 4

D 3 and 4

(c) Predict, without calculation, which reaction has the most negative value for  $\Delta S_{\text{system}}$ .

(1)

A 1

B 2

C 3

D 4

(Total for Question 11 = 3 marks)

TOTAL FOR SECTION A = 20 MARKS



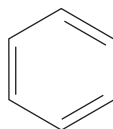


## SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

12 This question is about benzene and its compounds.

- (a) A Kekulé structure of benzene suggests the molecule consists of alternate double and single carbon to carbon bonds.



The standard enthalpy change of hydrogenation of a carbon to carbon double bond is  $-120 \text{ kJ mol}^{-1}$ .

- (i) Calculate the standard enthalpy change of hydrogenation of benzene for the Kekulé structure.

(1)

- \* (ii) The actual standard enthalpy change of hydrogenation of benzene is  $-208 \text{ kJ mol}^{-1}$ .

Use this information and your answer to (i), to calculate the difference in stability of benzene and the Kekulé structure.

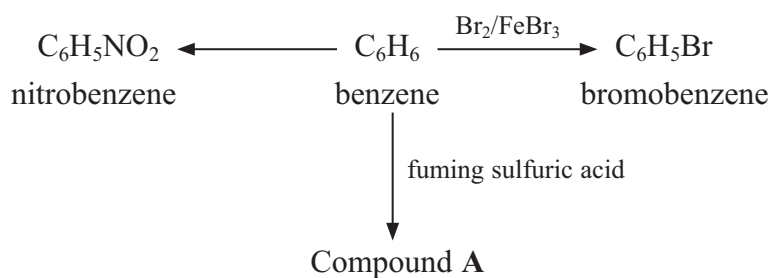
What does this tell us about the bonding in benzene?

Explain how this influences the type of chemical reactions that benzene undergoes.

(4)



(b) The flow chart shows some typical reactions of benzene.



(i) Identify the two chemicals needed to make nitrobenzene from benzene.

(2)

(ii) Classify the type of reagent required in all three reactions.

(1)

(iii) Show, with appropriate diagrams, the mechanism for the reaction of benzene with bromine, in the presence of iron(III) bromide, to form bromobenzene. Include the formation of the ion attacking the benzene molecule.

(4)



(iv) Give the structural formula and the name for compound A.

(2)

Name.....

(c) (i) Explain why phenol,  $C_6H_5OH$ , and methoxybenzene,  $C_6H_5OCH_3$ , are much more reactive than benzene with bromine.

(3)

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(ii) Write the equation for the reaction between phenol and bromine water. State symbols are **not** required.

(2)

(d) Diphenylmethane,  $(C_6H_5)_2CH_2$ , is often added to soap to create the pleasant smell of geraniums. It can be made from benzene and one other reagent, using a suitable catalyst.

Identify the other starting material and the catalyst by name or formula.

(2)

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(Total for Question 12 = 21 marks)

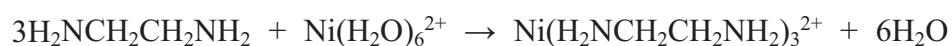


13 This question is about 1,2-diaminoethane,  $\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2$ .

- (a) Write an equation for the reaction between 1,2-diaminoethane and **excess** dilute hydrochloric acid.

(2)

- (b) When an aqueous solution of nickel(II) ions is added to 1,2-diaminoethane, a complex ion forms.



- (i) Suggest the colour of this complex.

(1)

- (ii) Without using the data booklet, suggest why the complex formed is more stable than  $\text{Ni}(\text{H}_2\text{O})_6^{2+}$  by considering the appropriate entropy change.

(2)

- (iii) This complex can exist as two structures, which are non-superimposable mirror images.

Suggest the physical property that would enable you to distinguish between these two structures.

(1)



(c) 1,2-diaminoethane reacts with butane-1,2-dioyl dichloride,  $\text{ClOCCH}_2\text{CH}_2\text{COCl}$ , to form a polymer.

(i) Draw a displayed formula for this polymer showing one repeat unit. (2)

(ii) Give, and explain, the name for this type of polymerization reaction. (2)

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\*(iii) State and explain the intermolecular forces between two of these polymer chains. (5)

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(Total for Question 13 = 15 marks)



14 An organic compound **X** contains carbon, hydrogen, oxygen and nitrogen.

0.132 g of **X** is burned completely in oxygen to produce 0.072 g of water, 0.176 g of carbon dioxide and 24.0 cm<sup>3</sup> of nitrogen.

[Molar volume of nitrogen under the conditions of the experiment = 24 000 cm<sup>3</sup> mol<sup>-1</sup>]

(a) Calculate the empirical formula of compound **X**.

(5)

(b) The molar mass of **X** is 132 g mol<sup>-1</sup>. Deduce the molecular formula for **X**. Show how you arrived at your answer.

(1)



- (c) When **X** is refluxed with concentrated hydrochloric acid for several hours, cooled and neutralized, there is only one organic product, **Y**, which has the molecular formula  $C_2H_5O_2N$ .

One mole of **Y** will react with either one mole of hydrochloric acid or one mole of sodium hydroxide solution.

When **Y** is sprayed with a solution of ninhydrin and heated, a purple colour is observed.

- (i) Use **all** the information above to deduce the functional groups present in **Y** and to classify the type of compound it is. Justify your answer.

(3)

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- (ii) Deduce the displayed formula for **Y** and give its name.

(2)

Name of **Y**.....

- (iii) Hence give the structural formula for **X**.

(1)

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(Total for Question 14 = 12 marks)

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**TOTAL FOR SECTION B = 48 MARKS**



## SECTION C

Answer ALL the questions. Write your answers in the spaces provided.

15

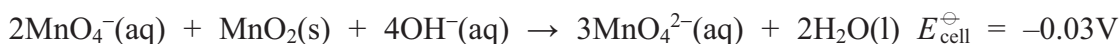
**Some unusual oxidation states**

You will probably have seen manganese in its common oxidation states during your study of chemistry. Manganese(II) sulfate occurs as very pale pink crystals in the hydrated form. Manganese(IV) oxide is a black powder which is often used as a catalyst. Potassium manganate(VII) occurs as very dark purple crystals and forms a purple aqueous solution, which is a powerful oxidizing agent.

You are less likely to have seen compounds containing the other oxidation states of manganese, which are manganese(VI), manganese(V), manganese(III) and manganese(I). However, compounds containing each of these four oxidation states can be prepared.

**Manganese(VI)**

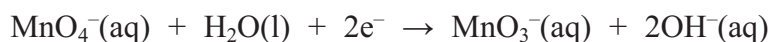
Manganese(VI), in  $\text{MnO}_4^{2-}$ , can be prepared in a reverse disproportionation reaction, by reacting manganate(VII) ions with manganese(IV) oxide in alkali.

**Equation 1**

The reaction is not thermodynamically favourable under standard conditions. However, the  $E_{\text{cell}}$  value can be made positive by increasing the concentration of hydroxide ions so that green manganate(VI) ions form.

**Manganese(V)**

Manganese(V) can be formed by adding a little potassium manganate(VII) to very concentrated ( $12 \text{ mol dm}^{-3}$ ) aqueous sodium hydroxide. The solution slowly becomes blue as manganate(V) ions,  $\text{MnO}_3^-(\text{aq})$ , form. The ionic half-equations are:

**Equation 2****Equation 3****Manganese(III)**

A deep red solution containing manganese(III) ions is formed by the oxidation of manganese(II) hydroxide by potassium manganate(VII) in acid solution. The ionic equation for the reaction is:

**Equation 4****Manganese(I)**

Manganese(I) ions are not stable in aqueous solution, but do form stable complex ions. They can be made by reducing hexacyanomanganate(II) ions,  $\text{Mn}(\text{CN})_6^{4-}$ , to hexacyanomanganate(I) ions,  $\text{Mn}(\text{CN})_6^{5-}$ .





(a) (i) Give the formula of manganese(IV) oxide. (1)

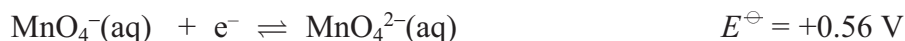
(ii) How do catalysts speed up reactions? (2)

(iii) Explain how transition metal ions can act as homogeneous catalysts. (2)

(b) (i) Suggest why the preparation of manganate(VI) ions,  $\text{MnO}_4^{2-}$ , in **equation 1**, may be described as a reverse disproportionation reaction by considering the relevant oxidation states. (2)



(ii) The two half-equations which are combined to form **equation 1** are



Explain, by reference to these half-equations, why increasing the concentration of hydroxide ions alters the electrode potential to make the preparation of manganate(VI) ions more likely.

(3)

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(c) Use equations 2 and 3 to answer the following questions.

(i) Identify the gas formed in the preparation of manganate(V) ions.

(1)

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(ii) By appropriately combining these two equations, write the ionic equation for the formation of manganate(V) ions from manganate(VII) ions.

(2)

(iii) Identify the main hazard and state how you would minimize the associated risk in this preparation of manganate(V) ions.

(2)

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(d) Identify the reagents you would use to make manganese(II) hydroxide for the preparation of manganese(III) ions.

(2)

(e) (i) Draw a dot and cross diagram to show the electron arrangement in the cyanide ion,  $\text{CN}^-$ .

(2)

(ii) Explain how the cyanide ion acts as a ligand.

(2)

(iii) Suggest the name of the shape of the hexacyanomanganate(I) ion.

(1)

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(Total for Question 15 = 22 marks)

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**TOTAL FOR SECTION C = 22 MARKS**

**TOTAL FOR PAPER = 90 MARKS**



